Chemistry 115 Name

Dr. Cary Willard

Exam 3b May 1, 2014

 Multiple Choice (30 points)

 Page 5 (24 points)

 Page 6 (16 points)

 Page 7 (17 points)

 Page 8 (13 points)

 Total (100 points)

 Percent

All work must be shown to receive credit. Give all answers to the correct number of significant figures

PV=nRT

Avogadros number = 6.022 x 1023 /mol

Ideal gas constant = 0.0821 L atm/mol K

 = 62.4 L torr/mol K

1 atm = 760 torr = 760 mm Hg = 101.3 kPa = 14.7 psi

Grossmont College

Periodic Table

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|  IA |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | VIIA | NOBLE GASES |
| 1**H**1.008 | IIA |  |  |  |  |  |  |  |  |  |  | IIIA | IVA | VA | VIA | 1**H**1.008 | 2**He**4.002 |
| 3**Li**6.941 | 4**Be**9.012 |  |  |  |  |  |  |  |  |  |  | 5**B**10.81 | 6**C**12.01 | 7**N**14.01 | 8**O**16.00 | 9**F**19.00 | 10**Ne**20.18 |
| 11**Na**23.00 | 12**Mg**24.30 | IIIB | IVB | VB | VIB | VIIB |  VIII VIII VIII | IB | IIB | 13**Al**27.00 | 14**Si**28.09 | 15**P**30.97 | 16**S**32.06 | 17**Cl**35.45 | 18**Ar**39.95 |
| 19**K**39.10 | 20**Ca**40.08 | 21**Sc**44.96 | 22**Ti**47.90 | 23**V**50.94 | 24**Cr**52.00 | 25**Mn**54.94 | 26**Fe**55.85 | 27**Co**58.93 | 28**Ni**58.70 | 29**Cu**63.55 | 30**Zn**65.38 | 31**Ga**69.72 | 32**Ge**72.59 | 33**As**74.92 | 34**Se**78.96 | 35**Br**79.90 | 36**Kr**83.80 |
| 37**Rb**85.47 | 38**Sr**87.62 | 39**Y**88.91 | 40**Zr**91.22 | 41**Nb**92.91 | 42**Mo**95.94 | 43**Tc**(99) | 44**Ru**101.1 | 45**Rh**102.9 | 46**Pd**106.4 | 47**Ag**107.9 | 48**Cd**112.4 | 49**In**114.8 | 50**Sn**118.7 | 51**Sb**121.8 | 52**Te**127.6 | 53**I**126.9 | 54**Xe**131.3 |
| 55**Cs**132.9 | 56**Ba**137.3 | 57**La**138.9 | 72**Hf**178.5 | 73**Ta**180.9 | 74**W**183.9 | 75**Re**186.2 | 76**Os**190.2 | 77**Ir**192.2 | 78**Pt**195.1 | 79**Au**197.0 | 80**Hg**200.6 | 81**Tl**204.4 | 82**Pb**207.2 | 83**Bi**209.0 | 84**Po**(209) | 85**At**(210) | 86**Rn**(222) |
| 87**Fr**(223) | 88**Ra**226.0 | 89**Ac**227.0 | 104**Rf**(261) | 105**Db**(262) | 106**Sg**(263) | 107**Bh**(262) | 108**Hs**(265) | 109**Mt**(266) | 110**??**(269) |  |  |  |  |  |  |  |  |

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| 58**Ce**140.1 | 59**Pr**140.9 | 60**Nd**144.2 | 61**Pm**(147) | 62**Sm**150.4 | 63**Eu**152.0 | 64**Gd**157.3 | 65**Tb**158.9 | 66**Dy**162.5 | 67**Ho**164.9 | 68**Er**167.3 | 69**Tm**168.9 | 70**Yb**173.0 | 71**Lu**175.0 |
| 90**Th**232.0 | 91**Pa**231.0 | 92**U**238.0 | 93**Np**(237) | 94**Pu**(244) | 95**Am**(243) | 96**Cm**(247) | 97**Bk**(247) | 98**Cf**(251) | 99**Es**(252) | 100**Fm**(257) | 101**Md**(258) | 102**No**(259) | 103**Lr**(260) |

Lanthanide series

Actinide series

Part I – Multiple Choice (30 points)

1. How many orbitals are contained in the 2p sublevel?
	1. 3
	2. 1
	3. 2
	4. 4
2. How many valence electrons are present in the element with the following ground state electron configuration?

1s2 2s2 2p3

* 1. 2
	2. 5
	3. 3
	4. 7
1. Which element is a transition element?
	1. Ca
	2. Kr
	3. K
	4. Co
2. What is the number of valence electrons in a halogen?
	1. 8
	2. 5
	3. 2
	4. 7
3. The maximum number of electrons that can fit into a single d orbital is
	1. 10
	2. 2
	3. 6
	4. 5
4. A mixture of gases consists of helium at a partial pressure of 400. torr, neon at a partial pressure of 300. torr, and argon at a partial pressure of 200. torr. What is the total pressure of this mixture of gases?
	1. 300. torr
	2. 900. torr
	3. 760. torr
	4. 1000 torr
5. Atoms of the metallic elements generally form ions by
	1. Gaining electrons, forming positive ions
	2. Gaining electrons, forming negative ions
	3. Losing electrons, forming positive ions
	4. Losing electrons, forming negative ions
6. When potassium fluoride forms from a potassium atom and a fluorine atom
	1. A proton is transferred from the potassium atom to the fluorine atom
	2. A proton is transferred from the fluorine atom to the potassium atom
	3. An electron is transferred from the potassium atom to the fluorine atom
	4. An electron is transferred from the fluorine atom to the potassium atom
7. Which element forms an ion that is larger than its atom?
	1. Lithium
	2. Calcium
	3. Chromium
	4. Fluorine
8. A Cl-1 ion has an electron configuration similar to that of
	1. Krypton
	2. Xenon
	3. Neon
	4. Argon
9. Carbon dioxide is a nonpolar molecule because
	1. Oxygen is more electronegative than carbon
	2. The two oxygen atoms are bonded to the carbon atom
	3. The molecule has a linear shape with the carbon in the middle
	4. The carbon-oxygen bonds are polar covalent
10. Which consists of molecules that are relatively far apart?
	1. Ne(g)
	2. Fe(s)
	3. HOH(l)
	4. Hg(l)
11. As the temperature of a gas sample increases, the number of molecules and volume remaining constant, the pressure exerted by the gas
	1. Increases
	2. Decreases
	3. Remains the same
12. The volume of a gas must always decrease when
	1. Temperature increases and pressure increases
	2. Temperature increases and pressure decreases
	3. Temperature decreases and pressure increases
	4. Temperature decreases and pressure decreases
13. At STP, which sample contains twice the number of molecules found in 4.00 L of hydrogen gas?
	1. 8.00 L of He
	2. 1.00 L of He
	3. 2.00 L of He
	4. 4.00 L of He

Part 2 – Problems and Short Answer (70 points)

1. (4 points) Write the complete electron configuration for an atom of aluminum. How many valence electrons does aluminum have?
2. (4 points) Write the shorthand electron configuration for an atom of paladium (Pd) as predicted by the periodic table.
3. (4 points) How is the line spectra of an atom produced? (Answer this on an atomic level)
4. (4 points) Explain how an ionic bond differs from a covalent bond.
5. (8 points) Choose the element/ion for each of the following:
	1. Which element has the higher ionization energy, Na or Si?
	2. Which element has the larger radius, Si or Ba?
	3. Which element is more electronegative, Mg or O?
	4. Which is larger, a sodium atom (Na) or a sodium ion (Na+)?
6. (8 points) Draw Lewis Electron Dot Structures for the following atoms/molecules.
	1. P
	2. C2H4 (Carbons are joined together and 2 hydrogens are attached to each carbon atom.)
	3. CH2O (Carbon is the central atom.)
7. (8 points) Predict the orbital or molecular geometry of the indicated atoms:

Orbital geometry Ca

Orbital geometry Cb

Molecular geometry Sc

Molecular geometry Nd

1. (5 points) Draw Lewis Electron Dot Structures for the nitrite ion (NO3−). Include reasonable resonance structures. (nitrogen is the central atom)
2. (8 points) A 35.0 L flask is filled with methane gas (CH4) at a pressure of 5.28 atm and a temperature of 42oC.
	1. How many moles of methane are contained in the flask?
	2. What is the pressure (in torr) of methane in the flask?
	3. If the temperature of the methane were raised to 79oC, what would the new pressure of the methane be?
3. (4 points) Hexane gas burns in the presence of oxygen gas to produce water and carbon dioxide.

2 C6H14 (g) + 19 O2(g) 🡪 14 H2O(g) + 12 CO2(g)

How many L of water will be produced if 5.00 L of hexane are burned in the presence of excess oxygen gas if the pressure and temperature are held constant?

1. (5 points) What is the density of a sample of carbon tetrabromide gas (CBr4) at a pressure of 4.67 atm and a temperature of 55oC?
2. (8 points) Given the following data

A sample of manganese is placed into a solution of nickel(II) chloride and the manganese dissolves. A precipitate of elemental nickel is observed to form. Another sample of manganese is placed into a solution of calcium chloride and no reaction occurs.

* 1. Write the equation for the reaction that takes place between nickel(II) chloride and manganese. (manganese will form an ion with a +2 charge)
	2. Which element is more active, nickel or manganese?
	3. Which element is more active, manganese or calcium?
	4. Would you expect a reaction to occur if you placed a sample of nickel metal into a solution of calcium chloride? Why or why not? (Hint: include the overall activity series in your answer)